	Inherent Qualities			Old Equation	New Equation	New Equation	New Equation reduced to common factor		Prime Numbers	Relative Abundance
	Vertexes	Faces	Edges	V + F = E + 2	✱ + F = E	÷ 2 ●				✱ + F = E
Tetrahedron	4	4	6	4 + 4 = 6 + 2	2 + 4 = 6	1 + 2 = 3	1 (1+2=3)		1	
Octahedron	6	8	12	6 + 8 =12 + 2	4 + 8 =12	2 + 4 = 6	2 (1+2=3)		2	
Cube	8	12	18	8 +12 =18 + 2	6 +12 =18	3 + 6 = 9	3 (1+2=3)	$\rho^2 \sim \nu^2 \cdot$	3	· ● · [1 + 2 = 3] + ◕
Icosahedron	12	20	30	12+20 =30 + 2	10+20 =30	5 +10 =15	5 (1+2=3)		5	
Vect. Equilib.	12	20	30	12+20 =30 + 2	10+20=30	5 +10 =15	5 (1+2=3)		5	

DEFINITIONS:

✱ Number of points (vertexes) other than those on poles = (V-2=✱) = non-polar vertexes.
◕ Polarity Constant that modifies all systems under consideration, additive twoness.
● Zonality Constant (Zone of Tunability), multiplying twoness.
V Number of vertexes.
F Number of faces.
E Number of edges.
ν Frequency - Modular breakdown.
ρ Wave length.

Gibbs' Phase Rule: F = C + 2 - P

where: F = Degrees of Freedom, i.e. number of variables.
C = Number of Chemical Components.
P = Phases of the System.
2 = Constant.

The phase rule is an equation for determining the number of possible degrees of freedom (variables) that can be given arbitrary values in a system in equilibrium without upsetting the equilibrium. For example in a system consistiing of ice, water. and water vapor, there are three phases: vapor, liquid, and crystalline; and one component: water. Therefore: F = 0. The three phases of water can coexist in equilibrium at a fixed temperature and pressure only, there are no degrees of freedom.

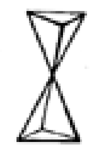
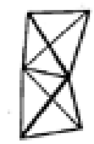
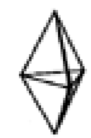
	Single-bonded	Double-bonded	Triple-bonded
Equivalents			
Phase	gas	liquid	crystalline
Bonds (vertexial)	single	double	triple
Connection	pin	hinge	fix
Inherent Qualities	vertex	edge	face

Fig. 1054.40